

concentration =

$$\text{Molarity (M)} = \frac{\text{moles of solute}}{\text{volume of solute}} = \left(\frac{\text{mol}}{\text{L}} \right) = M$$

3 M NaOH = "3 molar sodium hydroxide"

Dilution -

$$\underbrace{\text{Molarity (volume)}}_{\text{before}} = \underbrace{\text{Molarity (volume)}}_{\text{after}}$$

moles before = moles after

$$M_B V_B = M_A V_A$$

→ I want 25 mL of 2 M NaOH (after). My original solution is 5.5 M NaOH (before). How many mL of NaOH should