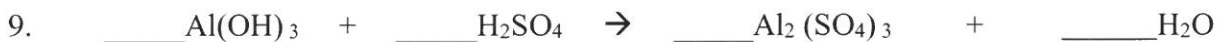
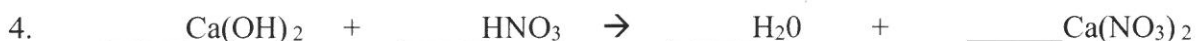
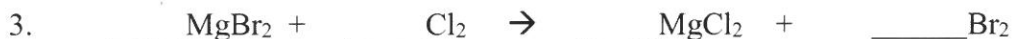


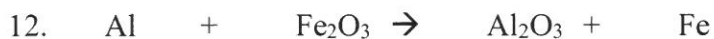
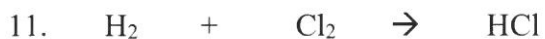
Name _____
Teacher _____
Class _____ Block ____
Date _____

Balancing Equations WS 1

Add the correct coefficient (number before the formula), as needed to the following equations.



Balance each of the following equations by adding the correct coefficients.



Name _____
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Class _____ Block _____
Date _____

Balancing Equations WS 2

Give the correct formulas for the following from the word equations and balance the chemical equations.

1. potassium bromide + barium iodide \rightarrow potassium iodide + barium bromide
2. copper (I) sulfide + oxygen \rightarrow copper (I) oxide + sulfur dioxide
3. aluminum + oxygen \rightarrow aluminum oxide
4. calcium hydroxide + ammonium sulfate \rightarrow calcium sulfate + ammonia + water
5. ethane (C₂H₆) + oxygen \rightarrow carbon dioxide + water
6. hydrochloric acid + magnesium hydroxide \rightarrow magnesium chloride + water
7. nitric acid + calcium hydroxide \rightarrow calcium nitrate + water
8. sodium chloride + hydrogen sulfate + manganese (IV) oxide \rightarrow manganese (II) sulfate + sodium sulfate + chlorine + water

Reaction Products Worksheet

For each of the following reactions, determine what the products of each reaction will be. When you have predicted the products, balance the equation and use a table of solubility products to determine which of the products (if any) will precipitate. Assume all reactions take place in water.

THERE IS ONE REACTION THAT WILL NOT TAKE PLACE BECAUSE THE ORIGINAL REACTANTS ARE NOT SOLUBLE IN WATER. It is one of the last 3 equations.

