Name _____
Teacher _____
Class _____ Block ___
Date _____

Balancing Equations WS 1

Add the correct coefficient (number before the formula), as needed to the following equations.

1. _____Mg + _____O_2 → _____MgO

2. _____$KClO_3$ → _____KCl + _____O_2

3. _____$MgBr_2$ + _____Cl_2 → _____$MgCl_2$ + _____Br_2

4. _____$Ca(OH)_2$ + _____HNO_3 → _____H_20 + _____$Ca(NO_3)_2$

5. _____Al + _____HCl → _____$AlCl_3$ + _____H_2

6. _____H_2 + _____N_2 → _____NH_3

7. _____$AgNO_3$ + _____$NaCl$ → _____$AgCl$ + _____$NaNO_3$

8. _____K + _____H_2O → _____KOH + _____H_2

9. _____$Al(OH)_3$ + _____H_2SO_4 → _____$Al_2(SO_4)_3$ + _____H_2O

10. _____NaI + _____Cl_2 → _____$NaCl$ + _____I_2

Balance each of the following equations by adding the correct coefficients.

11. H_2 + Cl_2 → HCl

12. Al + Fe_2O_3 → Al_2O_3 + Fe

13. $Pb(C_2H_3O_2)_2$ + H_2S → PbS + $HC_2H_3O_2$

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Balancing Equations WS 2

Give the correct formulas for the following from the word equations and balance the chemical equations.

1. potassium bromide + barium iodide → potassium iodide + barium bromide

2. copper (I) sulfide + oxygen → copper (I) oxide + sulfur dioxide

3. aluminum + oxygen → aluminum oxide

4. calcium hydroxide + ammonium sulfate → calcium sulfate + ammonia + water

5. ethane (C_2H_6) + oxygen → carbon dioxide + water

6. hydrochloric acid + magnesium hydroxide → magnesium chloride + water

7. nitric acid + calcium hydroxide → calcium nitrate + water

8. sodium chloride + hydrogen sulfate + manganese (IV) oxide → manganese (II) sulfate + sodium sulfate + chlorine + water

Reaction Products Worksheet

For each of the following reactions, determine what the products of each reaction will be. When you have predicted the products, balance the equation and use a table of solubility products to determine which of the products (if any) will precipitate. Assume all reactions take place in water.
THERE IS ONE REACTION THAT WILL NOT TAKE PLACE BECAUSE THE ORIGINAL REACTANTS ARE NOT SOLUBLE IN WATER. It is one of the last 3 equations.

1) ____ $Ca(OH)_2$ + ____ HF →

2) ____ $Pb(NO_3)_2$ + ____ K_2CrO_4 →

3) ____ $NaC_2H_3O_2$ + ____ H_2SO_4 →

4) ____ $Cu(OH)_2$ + ____ H_3PO_4 →

5) ____ $AgNO_3$ + ____ Na_2CO_3 →

6) ____ Zn + ____ H_2CO_3 →

7) ____ $Pb(OH)_2$ + ____ Hg_2S →