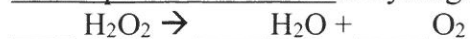


Name \_\_\_\_\_  
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Class \_\_\_\_\_ Block \_\_\_\_  
Date \_\_\_\_\_

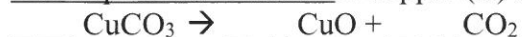
### Balancing Types of Reactions

Balance the following equations using **atom inventory** and **coefficients**. If it is balanced, write "balanced."

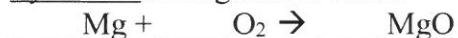
1. Decomposition Reaction of hydrogen peroxide.



2. Decomposition Reaction of copper (II) carbonate.



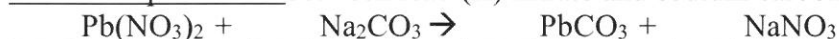
3. Synthesis of magnesium oxide.



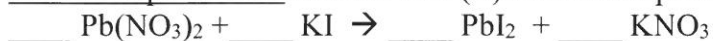
4. Synthesis of copper (II) oxide.



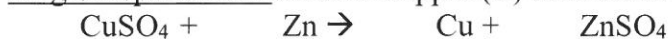
5. Double Replacement between lead (II) nitrate and sodium carbonate.



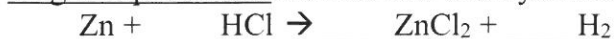
6. Double Replacement between lead (II) nitrate and potassium iodide.



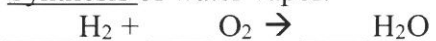
7. Single Replacement between copper (II) sulfate and zinc.



8. Single Replacement between zinc and hydrochloric acid.



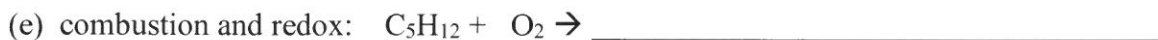
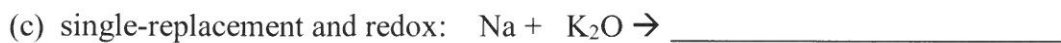
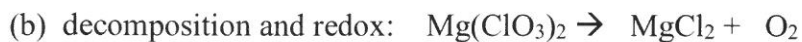
9. Synthesis of water vapor.



Name \_\_\_\_\_  
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Class \_\_\_\_\_ Block \_\_\_\_  
Date \_\_\_\_\_

### Types of Reactions WS

1. Complete and balance each of the following reactions identified by the type:



2. Classify each of the following reactions as synthesis (S), decomposition (D), single-replacement (SR), double-replacement (DR), combustion (C), redox (R), precipitation (P), or acid/base (A/B) by circling the correct answer:



3. Complete and **balance** each of the following equations and identify each type of reaction.

