

28. Give the orbital notation for lithium.
29. Give the electron configuration for magnesium.
30. Give the noble gas configuration for lanthanum.

31. Give the orbital notation for boron.

32. Give the electron configuration for boron.

33. Write the molecular equation for the reaction that produces solid lead (II) iodide and aqueous potassium nitrate. Show all states of matter on all reactants and products. The equation must be balanced.



34. Write the complete ionic equation for the reaction that produces solid lead (II) iodide and aqueous potassium nitrate. Show all states of matter on all reactants and products. The equation must be balanced.