Nam	e Hook time
	S Block
_	her
Date	Stoichiometry and Titration
Shov	v all work for the following problems. 259. Figs
1.	How many moles of water are produced if 3.6 moles of hydrogen react with 3.6 moles of oxygen?
Limiting reactant:	
	ting reactant: 3.6 mol t_2 x $\frac{2 mol t_20}{2 mol t_20} = \frac{3.6 mol t_20}{120}$ $L.R. = t_2$ 3.6 mol t_2 x $\frac{2 mol t_20}{2 mol t_20} = 7.2 mol t_20$ How many moles of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
	21 1 6 2 molto = 7,2 mol 120
2. Limi	J. O mol O2 x
	How many moles of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
	ting reactant: 369Hz x InsolHz x 2mol HzO = 1.78 = 1.8 mol HzO
	ting reactant: $3(gH_2 \times \frac{1}{2}) = 1.78 = 1.8 \text{ mol } H_2O$ $3(gH_2 \times \frac{1}{2}) = 1.016 \text{ gHz}$ $3(gH_2 \times \frac{1}{2}) = 1.8 \text{ mol } H_2O$ $3(gH_2 \times \frac{1}{2}) = 1.8 \text{ mol } H_2O$ $3(gH_2 \times \frac{1}{2}) = 1.8 \text{ mol } H_2O$ $3(gH_2 \times \frac{1}{2}) = 1.8 \text{ mol } H_2O$ How many grams of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
3.	37, さらして How many grams of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
Limi	ting reactant: 2 molths 2 molths y 180169ths = 329 Hz0
	2 ollegh
	How many grams of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen? ting reactant: $ \frac{1}{3} \cdot \log H_2 \times \frac{1}{2} \cdot \log H_2 \times \frac{1}{3} \cdot \log H_2$
4.	How many grams of water are produced if 3.6 g of sodium hydroxide react with 3.6 g of hydrochloric acid?
Limi	ting reactant: NaOH + HC1 = NaC1 + H, O (Balanced) NaOH=2299+16+1008 HC1=1008+35.4
	3.6g NaOH x molNaOH × mol NaOH × Truck NaOH / Truck Ho T. 6g H2 O ILR = NaOH)
	3,69 HC1 x 1 molth 0 x 18.016 gth 0 = 1 89 H20
5.	What is the molarity of sodium hydroxide if 3.6 g is dissolved in 200.0 mL of water?

What is the molarity of hydrochloric acid if 3.6 g is dissolved in 200.0 mL of water?

What is the molarity of hydrochloric acid if it is diluted from 1 L of 12 M concentrated acid and you now have

What is the molarity of hydrochloric acid if 10 mL of it completely neutralizes with 45 mL of 2.0 M sodium

6.

7. 15 L?

8.

hydroxide?

Name Class	Block L.R. each time
	er
Date	Stoichiometry and Titration
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Show a	all work for the following problems.
1.	How many moles of water are produced if 3.6 moles of hydrogen react with 3.6 moles of oxygen?
Limiui	18 reactant: 3.6 mol Hz x I mol Oz = [1.8 mol Oz Oz is excess, the is limiting]
	3.6 mol Hz × 2 mol Hz 0 = 3.6 mol Hz 0
2.	How many moles of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
Limitii	3.6g0 ₂ × Insl Hz × Insl0 ₂ × 32.00g0 ₂ = 29g0 ₂ O ₂ is limiting 3.6g0 ₂ × Insl O ₂ × 2 mol H ₂ O = [0.23mol H ₂ O] 3.6g0 ₂ × Insl O ₂ × 2 mol H ₂ O = [0.23mol H ₂ O]
3.	$3.6gO_2$, $\frac{1molO_2}{32.00gO_2}$, $\frac{2molH_2O}{1molO_2} = 0.23molH_2O$ How many grams of water are produced if 3.6 g of hydrogen react with 3.6 g of oxygen?
Limitii	ng reactant:
	3.69 H2 x Imalla x Imaloz x 32,00902 = 29902 02 is limiting!
4	3.69 H2 x 1 mol 02 x 32,00 gDz = 29g 02 [02 is limiting] 3.69 D2 x 1 mol 02 x 2 mol H2 0 x 18.0169 H20 14.19 H20] 1 mol 02 x 1 mol 02 x 2 mol H20 x 18.0169 H20 14.19 H20]
4.	How many grams of water are produced it 3.5 g of sodium hydroxide react with 3.5 g of hydrochloric acid?
Limitii	18 reactant: 3.69 NaOHx 399989 NaOH + I mol HCI 36.4589HCI = 3.39 HCI NaOH is limiting
	3,69NaOH, I mol NaOH I mol NaOH I mol H2O = [1.69 H2O]
5	What is the molerity of sedium hydravide if 2.6 a is dissolved in 200.0 ml of water?

What is the molarity of hydrochloric acid if 3.6 g is dissolved in 200.0 mL of water?

What is the molarity of hydrochloric acid if it is diluted from 1 L of 12 M concentrated acid and you now have

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